

REVIEW OF THE ATOM

- Atoms are the building blocks of matter
- Elements are composed of only one type of atom
- Compounds are composed of two or more atoms chemically bonded together in a fixed proportion
- All atoms are made up of three subatomic particles

Particle	Relative Mass	Charge	Location
Proton	1	Positive	Nucleus
Neutron	1	Neutral	Nucleus
Electron	0	negative	Orbit around the nucleus

Determining the Number of Subatomic Particles in an Element:

Number of Protons = Atomic Number (Z)

Number of Neutrons = Mass Number (A) – Atomic Number (Z)

- The mass number is the rounded version of the atomic mass
Ex. *nitrogen has an atomic mass of 30.973761 so its mass number becomes 31*
- recall that the mass number is equal to the number of protons and neutrons

Number of Electrons = the number of protons in a NEUTRAL atom

Example:

